# CARBOHYDRATES: STRUCTURE AND FUNCTION

# Objectives

**To understand:**

- **The structure of carbohydrates of physiological significance**
- **The main role of carbohydrates in providing and storing of energy**
- **The structure and function of glycosaminoglycans**

# OVERVIEW

**Carbohydrates:**

**The most abundant organic molecules in nature**

**The empiric formula is $(CH_2O)n$, "hydrates of carbon"**

**Carbohydrates:**

**provide important part of energy in diet**

**Act as the storage form of energy in the body**

**are structural component of cell membranes**

# OVERVIEW

CONT'D

- Many diseases associated with disorders of carbohydrate metabolism including:

  Diabetes mellitus

  Galactosemia

  Glycogen storage diseases

  Lactose intolerance

# CLASSIFICATION

- **Monosaccharides**: Simple sugar
- **Disaccharides**: 2 monosaccharide units
- **Oligosaccharides**: 3-10 monosaccharide units
- **Polysaccharides**: more than 10 sugar units

  Homopolysaccharides &

# Monosaccharides

**Further classified based on:**

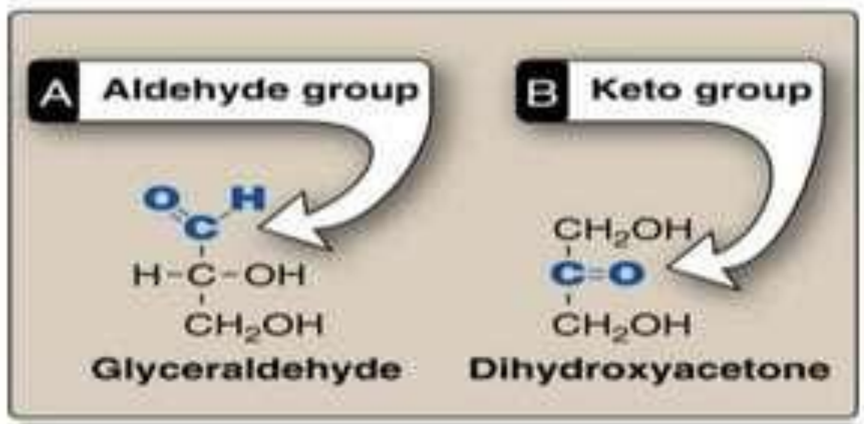
**1. No. of carbon atoms**

| Generic names | | Examples |
|---|---|---|
| 3 carbons: | trioses | Glyceraldehyde |
| 4 carbons: | tetroses | Erythrose |
| 5 carbons: | pentoses | Ribose |
| 6 carbons: | hexoses | Glucose |
| 7 carbons: | heptoses | Sedoheptulose |
| 9 carbons: | nonoses | Neuraminic acid |

**2. Functional sugar group:**

**Aldehyde group – aldoses**

**Keto group – ketoses**



# Monosaccharides

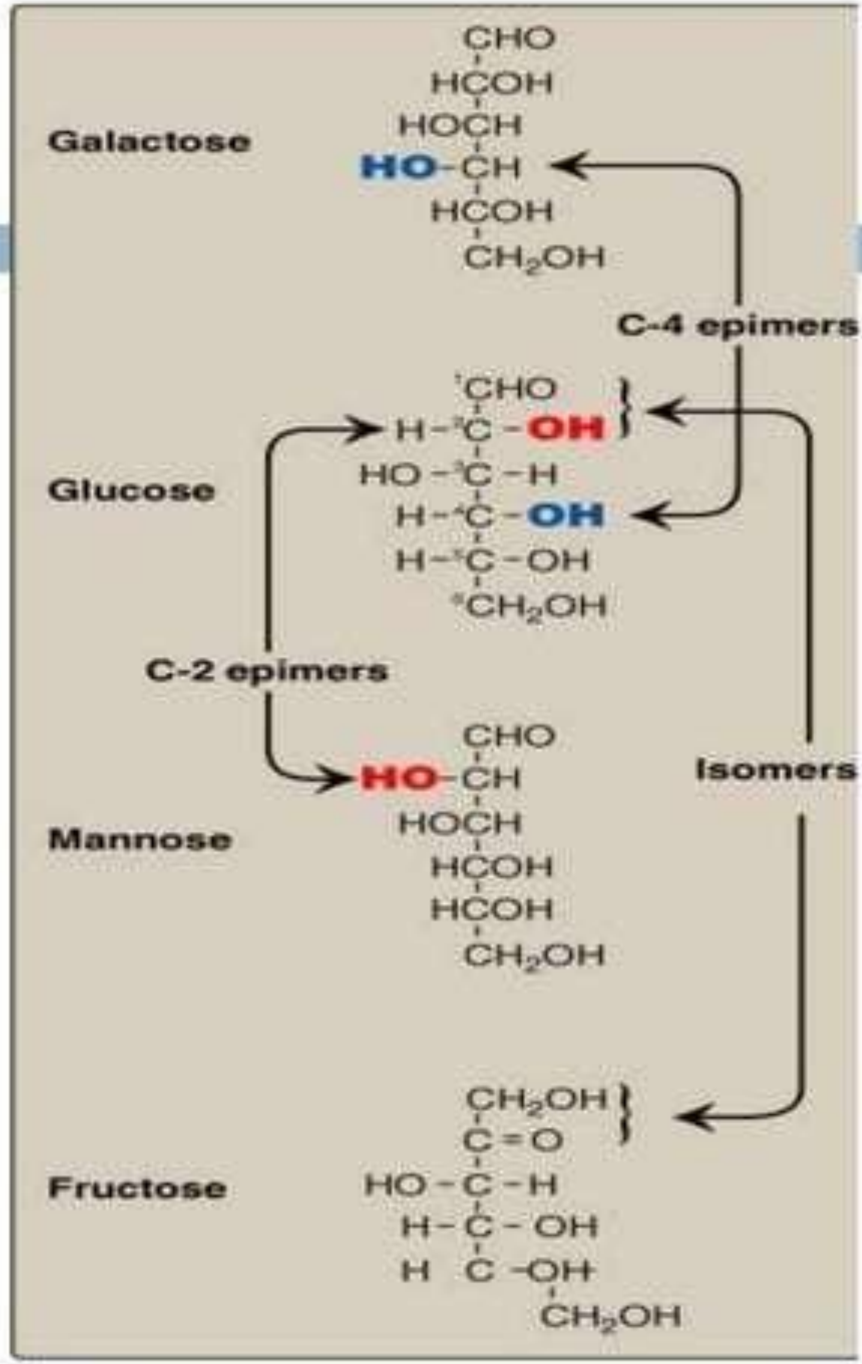
CONT'D

| | Aldose | Ketose |
|---|---|---|
| Triose | Glyceraldehyde | Dihydroxyacetone |
| Pentose | Ribose | Ribulose |
| Hexose | Glucose | Fructose |

# Isomerism

- Isomers

Compounds having same chemical formula but different structural formula



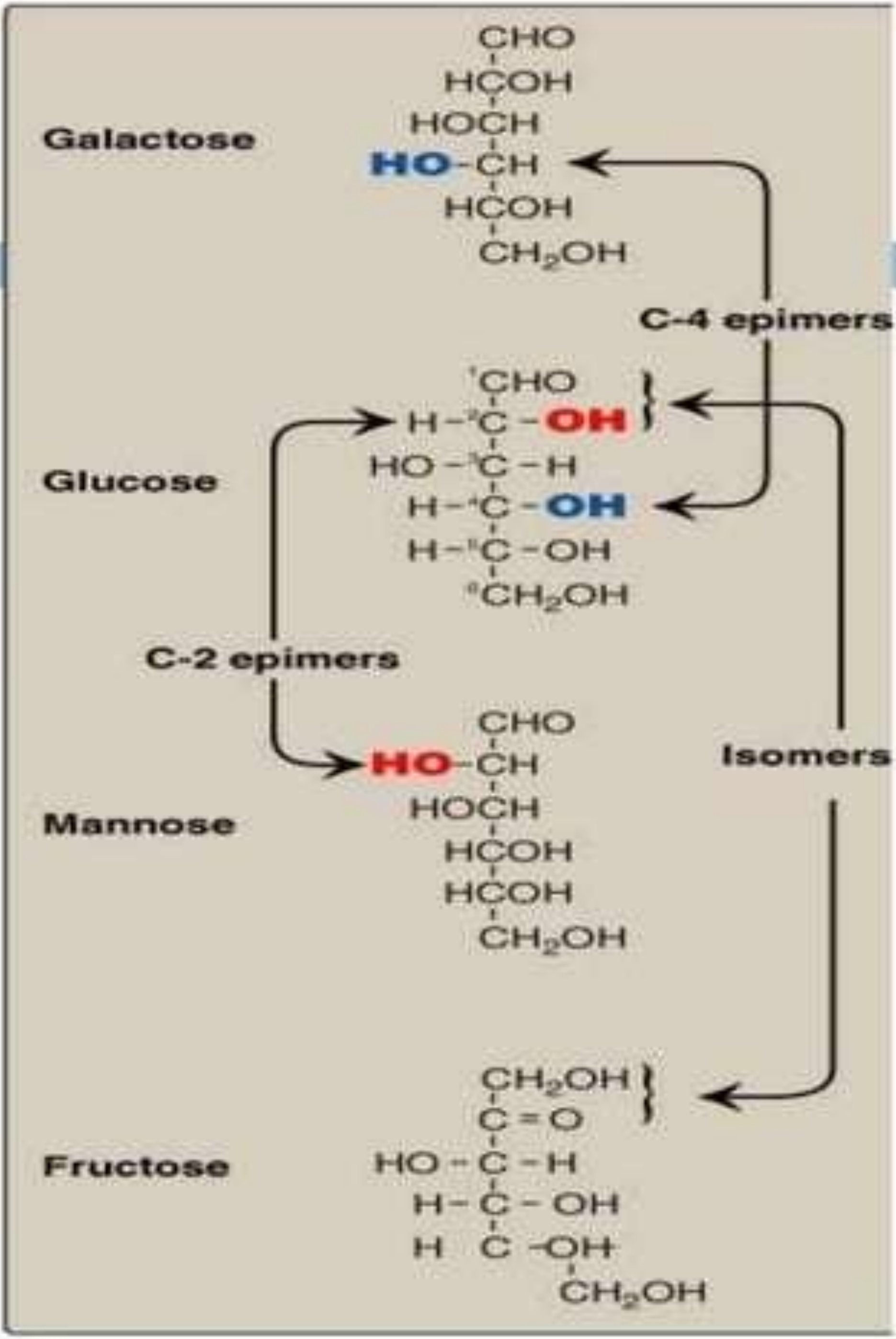
# Aldo-Keto Isomers

**Example:**

Glucose (Aldose)

and

Fructose (Ketose)



# Epimers

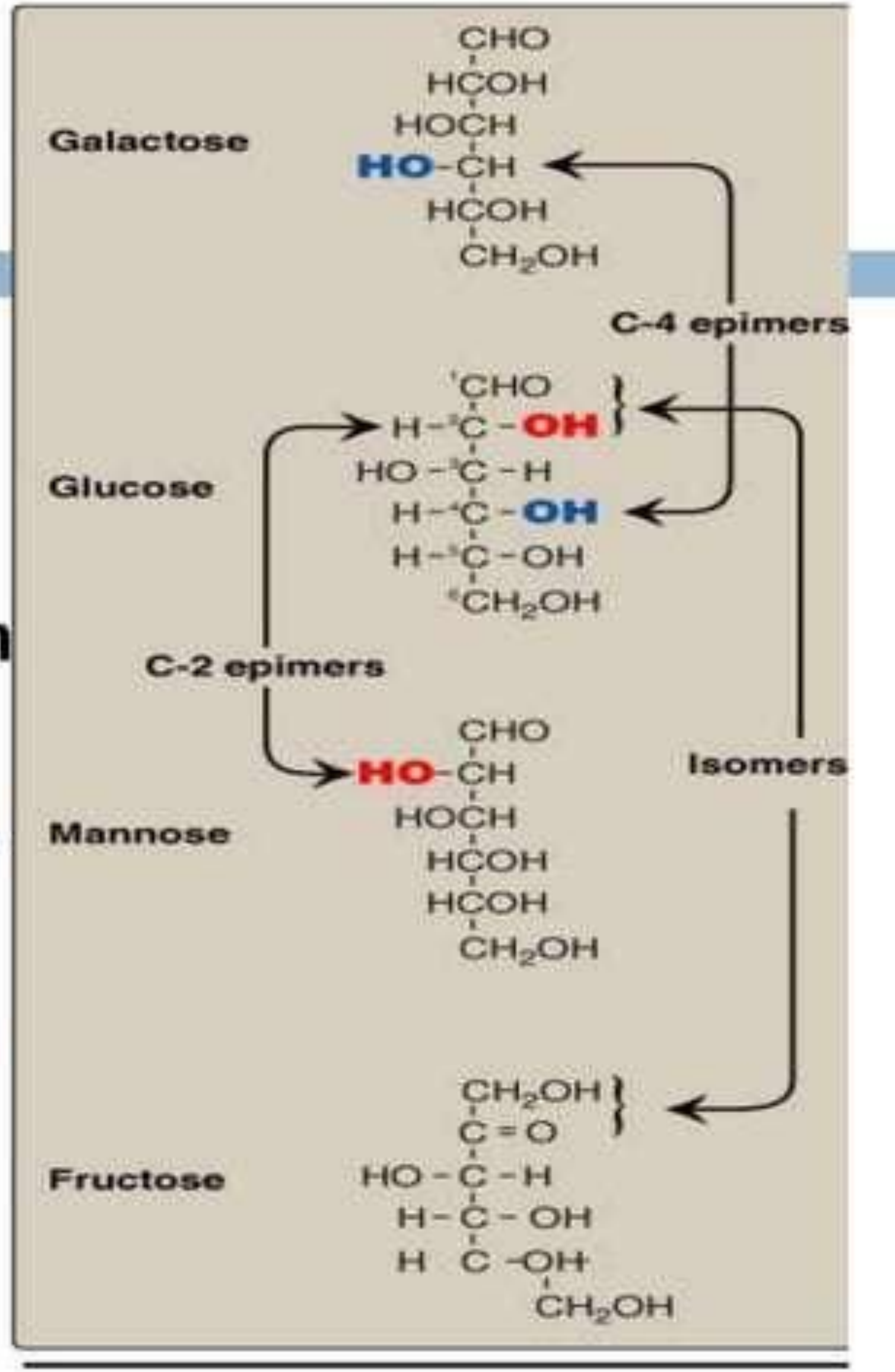
- Epimers

CHO dimers that differ in configuration around only one specific carbon atom

-Glucose and galactose, C4

-Glucose and Mannose, C2

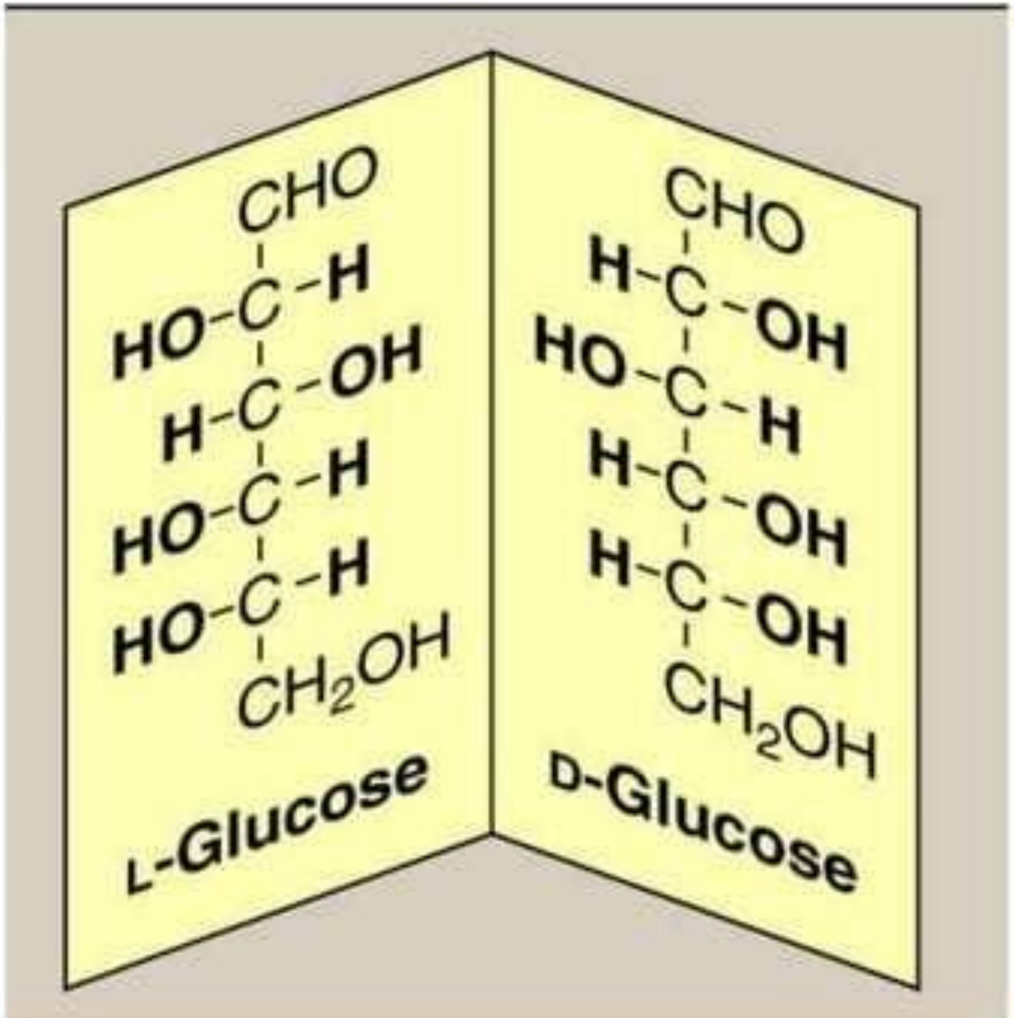
Galactose and mannose are not epimers



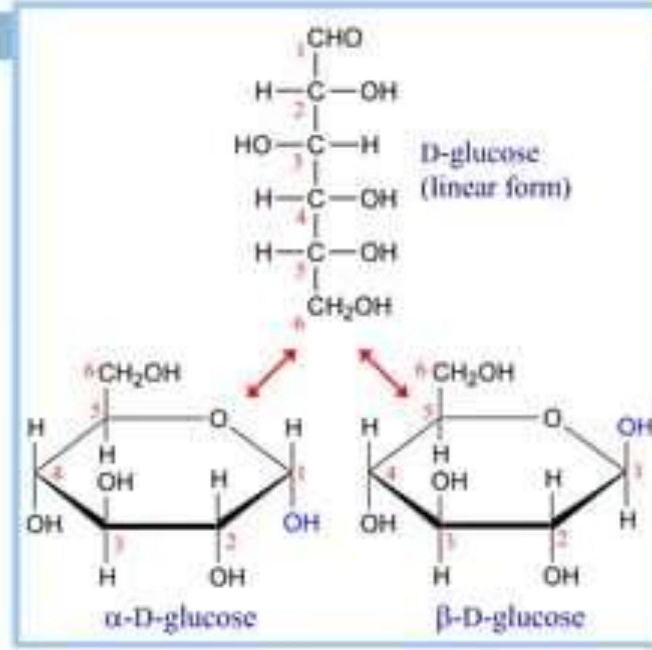
# Enantiomers (D- and L-Forms)

Structures that are **mirror images** of each other and are designated as D- and L- sugars based on the position of –OH grp on the **asymmetric carbon farthest from the carbonyl carbon**

Majority of sugars in humans are **D-sugars**



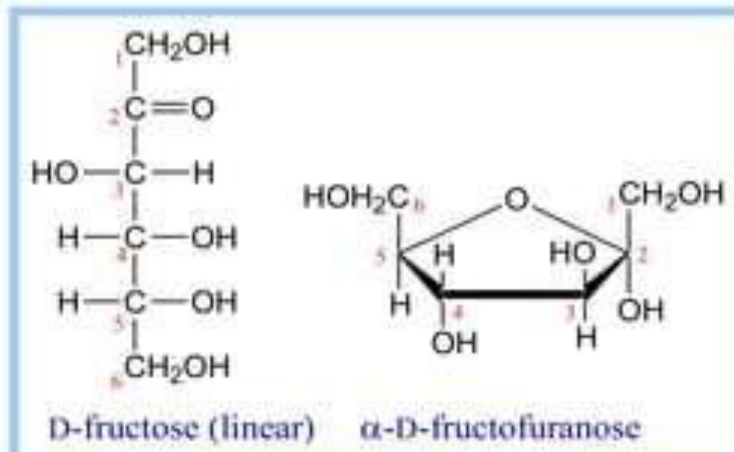
# α- and β-Forms




❑ **Cyclization of Monosaccharides**

Monosaccharides with 5 or more carbon are predominantly found in the ring form

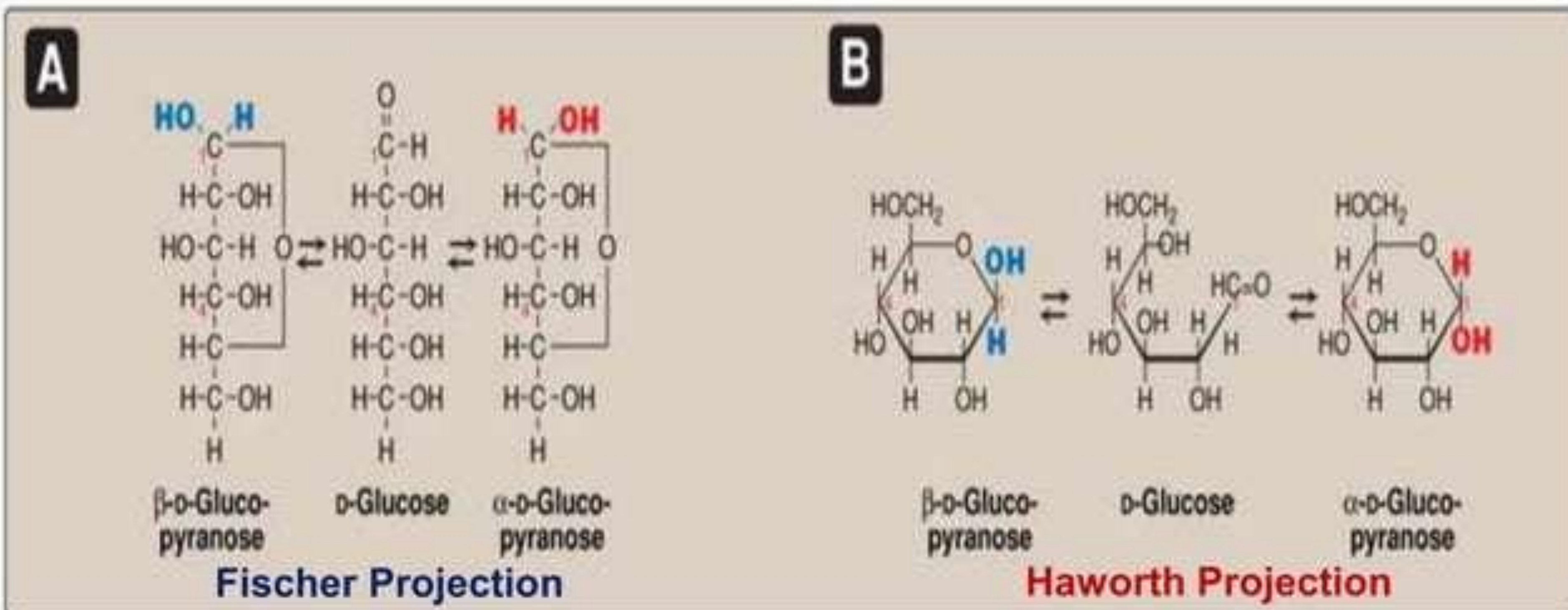
-The aldehyde or ketone grp reacts with the –OH grp on the same sugar




-Cyclization creates an **anomeric carbon** (former carbonyl carbon) generating the α and β configurations

# Mutarotation

In solution, the cyclic α and β anomers of a sugar are in equilibrium with each other, and can be interconverted spontaneously



# Sugar Isomers

1. **Aldo-keto**
2. **Epimers**
3. **D- and L-Forms**
4. **α- and β-anomers**

# Disaccharides

- Joining of 2 monosaccharides by O-glycosidic bond:

**Maltose** (α-1, 4)= glucose + glucose

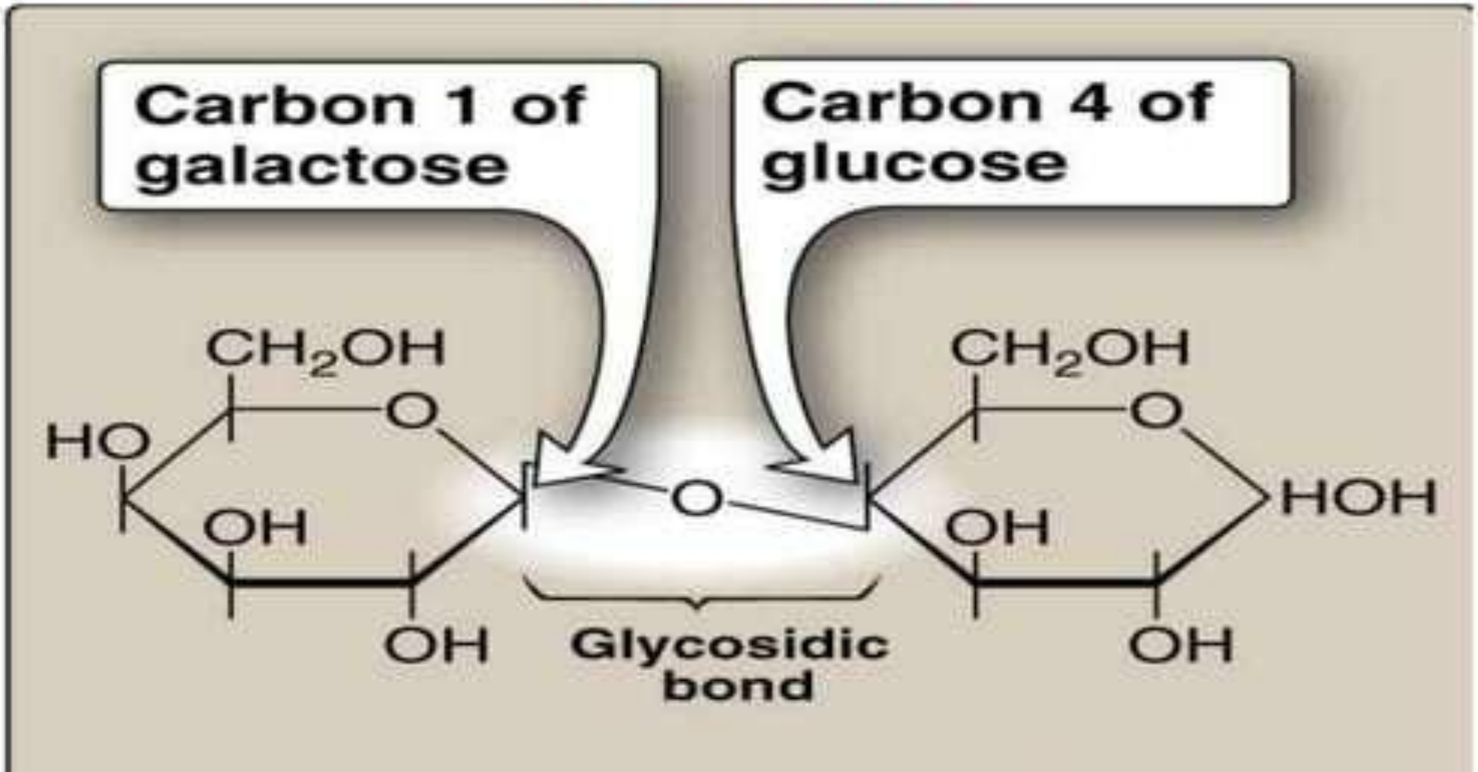
**Sucrose** (α-1,2) = glucose + fructose

**Lactose** (β-1,4) = glucose + galactose

# Disaccharides

CONT'D

## Lactose




Lactose: galactosyl-β(1→4)-glucose

# Polysaccharides

- **Homopolysaccharides:**

  **Branched:**

  Glycogen and starch (α-glycosidic polymer)

  **Unbranched:**

  Cellulose (β-glycosidic polymer)

- **Heteropolysaccharides:**

  e.g., glycosaminoglycans (GAGs)

# Reducing Sugars

- If the O on the anomeric C of a sugar is not attached to any other structure **(Free)**, that sugar can act as a reducing agent

- **R**educing sugars reduce chromogenic agents like Benedict's reagent or Fehling's solution to give a colored precipitate

- **U**rine is tested for the presence of reducing sugars using these colorimetric tests

# Reducing Sugars

CONT'D

- Examples:

  Monosaccharides

  Maltose and Lactose

  Sucrose is non-reducing, Why?

# Complex Carbohydrates

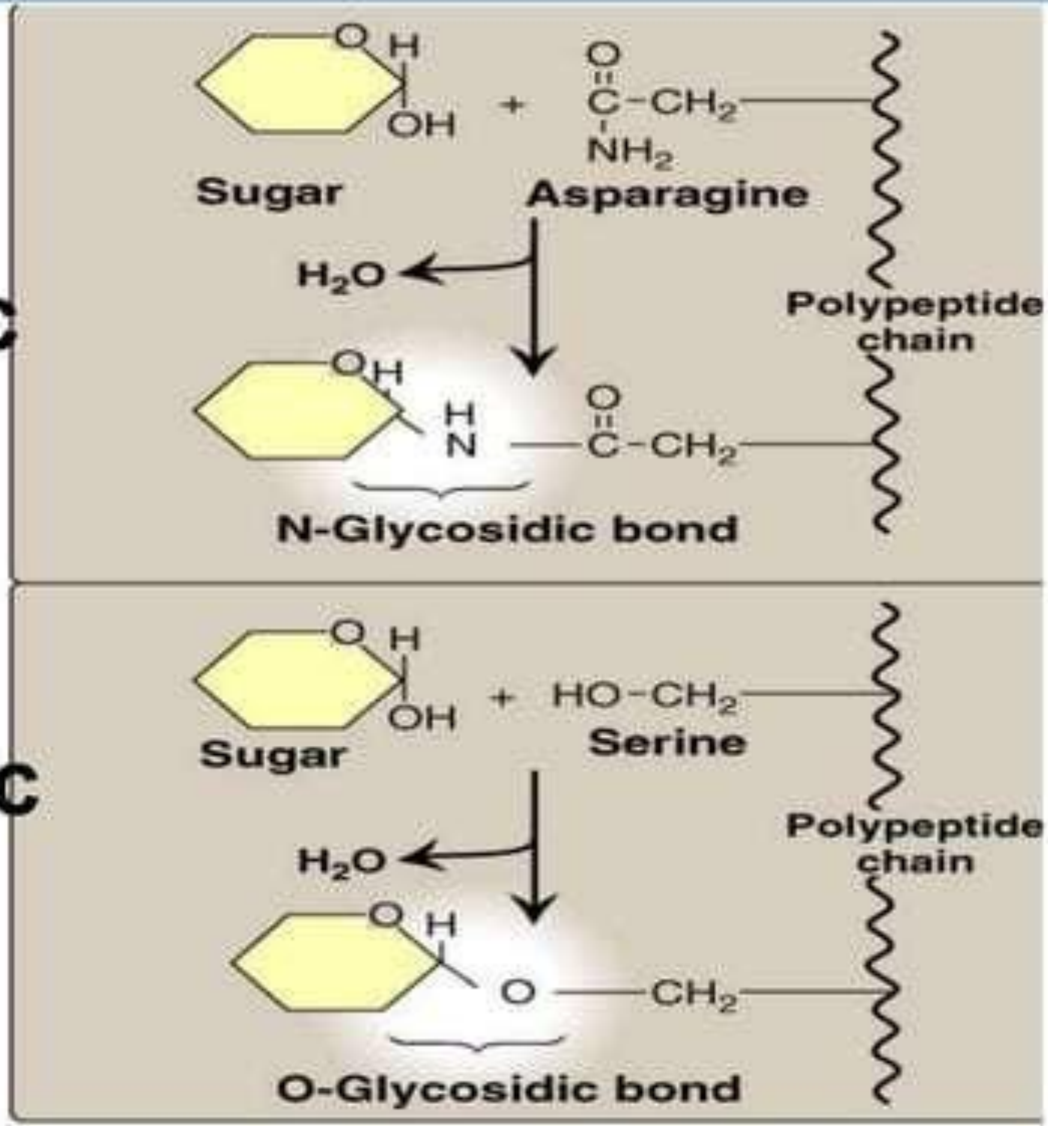
- Carbohydrates attached to non-carbohydrate structures by glycosidic bonds (O- or N-type) e.g.,

1. Purine and pyrimidine bases in nucleic acids
2. Bilirubin
3. Proteins in glycoproteins and proteoglycans

# Glycosidic Bonds

- N-Glycosidic

- O-Glycosidic



# Glycosaminoglycans (GAGs)

- Glycosaminoglycans (GAGs) are large complexes of **negatively** charged **hetero**polysaccharide chains
- are associated with a small amount of protein, forming **proteoglycans**, which consist of over 95 percent carbohydrate
- bind with large amounts of water, producing the gel-like matrix that forms body's ground substance
- The viscous, lubricating properties of mucous secretions also result from GAGs, which led to the original naming of these compounds as mucopolysaccharides

# Glycosaminoglycans (GAGs)

- **GAGs** are linear polymers of **repeating disaccharide** units

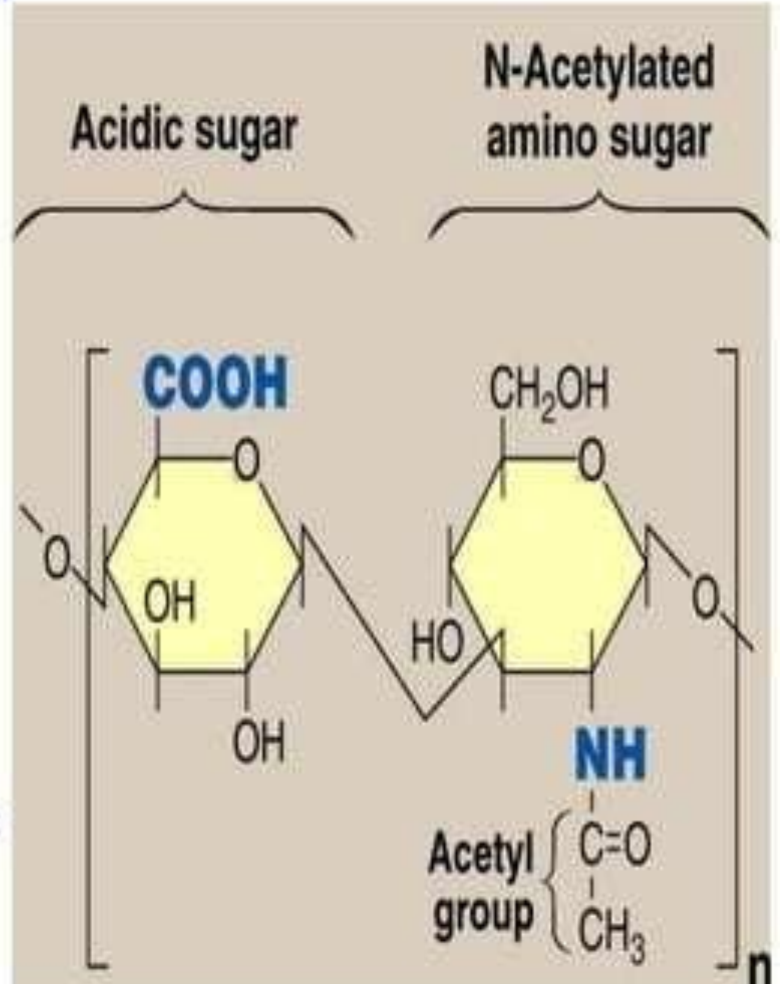
  **[acidic sugar-amino sugar]n**

- The amino sugar (usually sulfated) is either

  **D-glucosamine or D-galactosamine**

- The acidic sugar is either

  **D-glucuronic acid or L-iduronic acid**




- GAGs are strongly negatively-charged:

# Resilience of GAGs

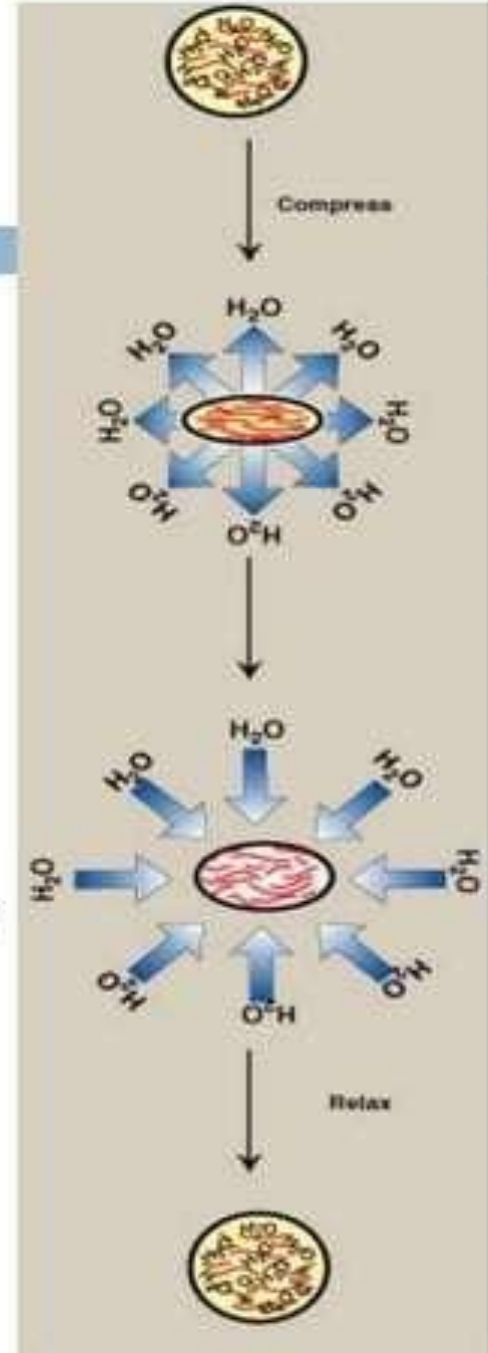
**Relationship between glycosaminoglycan structure and function**

- Because of negative charges, the GAG chains tend to be extended in solution and repel each other and when brought together, they "slip" past each other

This produces the <u>**"slippery" consistency of mucous secretions and synovial fluid**</u>

- When a solution of GAGs is compressed, the water is "squeezed out" and the GAGs are forced to occupy a smaller volume. When the compression is released, the GAGs spring back to their original, hydrated volume because of the repulsion of their negative charges

This property contributes to the <u>**resilience of**</u>



# Members of GAGs

**Examples of GAGs are:**

1. **Chondroitin sulfates: Most abundant GAG**
2. **Keratan sulfates:** Most heterogeneous GAGs
3. **Hyaluronic acid: Compared to other GAGs, it is unsulfated and not covalently attached to protein**
4. **Heparin: Unlike other GAGs,** Unlike other GAGs that are extracellular, heparin is **intracellular and serves as an anticoagulant**

# Take home Message

Structure and function of carbohydrates

- Mono-, Di-, and Poly-saccharides
- Sugar Isomers: Aldo-keto, epimers, D- and L-, α- and β-anomers
- Complex carbohydrates:

  e.g., Glycosaminoglycans and proteoglycans
- Structure and function of GAGs
- Examples of GAGs: chondroitin sulfate, keratin sulfate, hyaluronic acid and heparin

# Reference


- Lippincott's Illustrated reviews- Biochemistry, 6th Edition, pages- 83-86 and 157-159